

Zeolite supported Pd electrocatalyst nanoparticle characterization

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Highlights:

- Detail Pd cluster structure by the ex-situ Extended X-ray Adsorption Fine Structure (EXAFS) method.
- Evaluate Pd/Y-zeolite nanoparticle size under calcinations and reduction temperature.
- Measure Pd nanoparticle electrochemical activity in electrolyte solution.
- Discuss electrochemical reaction behaviour with respect to Pd particle size.
- Explore Pd particle location and Pd particle migration and electrochemical behaviour.

Abstract

A laboratory made 1.5 wt% Palladium (Pd) zeolite electrocatalyst is investigated using the Extended X-ray Adsorption Fine Structure (EXAFS) and Cyclic Voltammetry (CV) techniques to reveal Pd structure and resultant electrochemical performance. It was found that the

electrochemical activity of hydrogen charger transfer in the hydride region increased for electrocatalyst with large-size particles made at high temperature of 400 °C, compared to those with small-size particles calcined and reduced at temperature below 360 °C, at which no major discrepancies were observed between catalysts of different sizes. Furthermore, Pd particle location has played an important role to enhance electrocatalyst performance. The Pd atom tends to remain at small cages, i.e. zeolite sodalite cages or hexagonal prisms at calcinations and reduction temperatures below 360 °C. When temperature increases to about 400 °C, the majority Pd atoms tend to migrate from zeolite small cages to supercages and zeolite external structures with enhanced electrochemical performance.

Keywords: Pd nanoparticle mobility; Calcinations and reduction temperatures; Cyclic Voltammetry (CV); Ex-situ EXAFS; Electrochemical activity; Zeolite sodalite cages and hexagonal prisms.

1. Introduction

The hydrogen polymer electrolyte fuel cell (PEFC) has been extensively investigated over the past five decades, with most studies focused on electrochemical reactions including the hydrogen oxidation reaction (HOR) and oxygen reduction reaction (ORR) [1, 2]. While the Platinum (Pt) based electrocatalysts have been proved to be the best choice for HOR and ORR reactions due to their high activity and stability [2], the high (Pt) price has limited its wider industry applications. This largely stimulates researchers to search for alternative electrocatalysts with comparable performance but at much reduced cost. The Palladium (Pd) as one of the Pt-group metals has attracted significant attentions, owing to its relatively good HOR and ORR activities [3], and much low cost (only about one third of that of Pt). In general, the Pd-based catalyst has shown lower catalytic activity than those Pt electrocatalysts, but it

exhibits a high electrocatalytic activity for oxidation reaction using small organic molecules, e.g. methanol and ethanol due to a better resilience to CO poisoning [4]. Hence, Pd nanoparticle catalyst can be a potential good candidate as an electrocatalyst for fuel cell applications, if Pd catalytic activity and stability can be largely improved to be compatible with Pt electrochemical activity. This is the main motivation of this proposed study.

Recently, a lot of efforts have been made to focus on the development of carbon and graphene based Pd-nano electrocatalysts [5] using approaches such as surface modification, composition control [6], alloying [7] and spontaneous methods [8] for hydrogen fuel production in PEFCs. In addition, the ‘support promotion’ method is another approach employed to improve catalyst nanoparticle distribution for hydrogen fuel cell efficiency enhancement [2]. Zeolite is a type of porous amorphous material consisting of crystalline framework, formed by aluminosilicates structure comprised of SiO_4 and AlO_4 tetrahedral forms as the base units. These tetrahedral structures are interlinked with each other by sharing one oxygen atom. Each $[\text{AlO}_4]^-$ tetrahedron in the framework induces negative charges which can be compensated by a charge-balancing cations of Na^+ or H^+ ions situated in the zeolite channels or cavity. The extra framework of metal ions on zeolite active sites (also called Brønsted acidic sites) are mobilized to take part in an ion exchange process with the desired metal ions. The ‘X’ or ‘Y’ type of zeolite channel structure with exemplary surface area and cation exchange capacity makes it a desirable platform for adsorption of different species to produce fine metal particle distribution. Moreover, the larger surface area and the pore structure of zeolite enable to absorb the electrolyte and electrochemical reaction species, which can lead to a suitable condition for electrochemical measurement during diffusion-controlled mass transfer process [9].

Zeolite catalytic performance can be modified by pre-exchanging with ions to give the desired form and then lock the exchangeable metal ions into the position by calcinations and reduction process with fine-tuned temperature control to further promote catalytic activity of

metal particles [10]. For example, Spoto et al. investigated 4 wt% Pd loading on NH₄-Na-Y zeolite exchanged with aqueous solution of [Pd(NH₃)₄](NO₃)₂ under calcinations at 200 °C and reduction at 300 °C, respectively. The hydrogen atoms transiting through silicate rings are the possible mechanisms for hydrogen transport to and from the sodalite – cages due to hydrogen diffusion in active form through zeolite framework by spill-over effects that hydrogen can be adsorbed by Pd [11]. The proton is able to migrate to zeolite neighboring rings, facilitating the mobility of hydrogen atoms in the silicate framework. The penetration probability of hydrogen within sodalite cage and hexagonal prism can be further enhanced by the presence of Pd catalytic active site on zeolite. The large-size Pd nanoparticle is found to locate at the exterior of zeolite surface due to the migration of Pd atoms from zeolite internal structure to the external surface during reduction process [11]. Alternatively, the atomic hydrogen can diffuse into the zeolite sodalite cages by penetrating from hexagonal openings of the zeolite framework. Homeyer and Sachtler [12] have reported that Pd(NH₃)_x²⁺ ions release into zeolite supercages at calcinations temperature of 250 °C. The oxidation of Pd(NH₃)₄²⁺ ions were observed to produce Pd(NH₃)₂²⁺ ions in supercages and Pd(NH₃)₂²⁺ in sodalite cages [13]. Further temperature increase was found to result in producing Pd²⁺ ions at calcinations temperature over 300 °C. The Pd²⁺ ions were found to have a high mobility, migrating from zeolite external surface and supercages into zeolite small cages i.e. sodalite cages or hexagonal prism. Most Pd²⁺ ions were forced into sodalite cages at the calcinations temperature up to 500 °C.

The reduction process has shown that Pd atom is more favorable to remain in the zeolite supercages than the sodalite and hexagonal prism [11]. Bergeret et al. [14] and Kleine et al. [15] have found that the reduction can result in agglomeration of Pd atoms in the supercages with primary particle size less than 7 Å. These particles could extend over several cages at high temperature condition. Pd species were observed to migrate back into zeolite supercages at reduction temperature of 360 °C. Moreover, the mobility of Pd might be contributed by the

restoration of zeolite active sites (i.e. Brønsted acidic sites) via H₂ reduction [16]. H₂ can have affinity with proton to induce a ligand displacement with the –OH groups interacted to oxygen atoms in the sodalite cages and form the hydrogen bonded complex [11].

In a previous study, the in-situ EXAFS data analysis of Pd nanoparticle structure on Y zeolite in an electrochemical environment has predicted a first shell Pd-Pd distance of 2.73 Å with respect to a Pd first shell coordination number of 12 which give rise to an average radius of metal particle greater than 15 Å [11]. However, the average of Pd particle size is still smaller on zeolite than on carbon and those coated on microcellular graphitic carbon where Pd particles can be between 2.7 nm – 65.2 nm and 6.6 nm – 30.3 nm, respectively [17].

Pd also has a capability of bulk incorporation of hydrogen that can lead to the high solubility and mobility of hydrogen in Pd. This might be associated to the diffusion of atomic hydrogen produced by the disassociation of the H₂ molecule on the Pd metal particles due to hydrogen spillover ability on Pd within zeolite internal structure especially sodalite cages. Subsequently, the Pd sample activation procedure can be important to control the atomic hydrogen spillover efficiency on Pd [11, 18]. Recently, zeolite supported Pd material has attracted a significant interest for its applications to enhance the fuel cell efficiency.

Generally, the pathway of hydrogen oxidation reaction (HOR) can include three important steps that might occur either via Volmer-Fafel or Volmer-Heyrovsky reaction process or both [2, 19, 20]. In acid electrolyte solution, the activity of HOR on Pd or Pt surface follows a very identical trend of hydrogen evaluation reaction that follows three reaction steps, in which the first step involves the Volmer reaction to receive an electron by a proton, leading to an intermediate hydrogen atom to be adsorbed at the active sites [21], whilst the second step was given by Heyrovsky or Tafel reaction, in which hydrogen was produced by two adsorbed hydrogen atoms on active sites of electrocatalyst [2]. Moreover, the Pd was found to exhibit the lower kinetics of hydrogen evolution reaction (HER) than Pt. Thus, the different approaches [5,

6, 7, 8] have been explored to develop the cost effective Pd electrocatalyst system, such as, nanocomposites [6] supported on different porous medium materials.

The study presented in this paper will focus on the investigation of Pd sample activating procedure that might affect the electrochemical performance of hydrogen oxidation and oxygen reduction process on Pd active sites in the electrolyte solution. The electrocatalytic activity of 1.5 wt% Pd on zeolite made by ion exchange method under various calcinations and reduction temperatures was investigated using Cyclic Voltammetry (CV) technique. Pd particle size and its dispersion on zeolite will be characterized by the ex-situ extended X-ray adsorption fine structure (EXAFS) using pallet made by Pd zeolite catalyst and boron nitride. The study will contribute to the well understanding of Pd electrocatalyst efficiency enhanced by hydrothermal process under calcinations and reduction temperature control.

2. Experimental

2.1 Preparation of Pd/Y zeolite electrocatalysts

The ion exchange method has been employed to achieve 1.5 wt% Pd loading on Y zeolite using $(\text{NH}_3)_4\text{PdCl}_4$ [7,19, 22]. In this process, the $[\text{Pd}(\text{NH}_3)_4]^{4+}$ ligand was ion exchanged with Na^+ ions situated on the alumina tetrahedral arrangement of $[\text{AlO}_4]^-$ ions that led to a high dispersion of Pd on zeolite with uniform distributions [22, 23]. Each positive $[\text{Pd}(\text{NH}_3)_4]^{4+}$ ion can interact with four negative charged $[\text{AlO}_4]^-$ ions to introduce a charge balance. Then, further calcinations reaction can induce a Pd^{4+} ion which bridges with four $[\text{AlO}_4]^-$ ions after decomposition of ammonia ligand.

The ion exchange process was carried out at PH7 in a neutral solution. An appreciated quantity of $(\text{NH}_3)_4\text{PdCl}_4$ was introduced in 200 ml of triply distilled water to produce a Pd concentration of $0.004 \text{ mol dm}^{-3}$. The Pd salt was dissolved uniformly in an ultrasonic bath at ambient room temperature. The process of $[\text{Pd}(\text{NH}_3)_4]^{4+}$ ions exchange with zeolite was taken

place using a water jacketed reaction with water cycled around at 70 °C, in which Y zeolite powder was dispersed at a concentration of 1 g per 100 ml of triply distilled water. The Pd salt solution was then pumped into the reactor at a flow rate of 0.1 ml/min. The resultant Pd/Y zeolite samples were washed with a sufficient amount of triply distilled water until no Cl⁻ ions detected by 1 N AgNO₃ solution. The completion of Pd complex ion exchange was confirmed by UV spectrum.

2.2 Synthesis of 1.5 wt% Pd zeolite electrocatalyst

1.5 wt% Pd/Y zeolite electrocatalysts were made by calcinations and reduction processes to remove Pd coordinate ligand. The [Pd(NH₃)₄]⁴⁺ salt loaded Y zeolite samples were purged with argon gas at 150 °C to dry out the water moisture in a fluidized reactor. The calcinations of [Pd(NH₃)₄]⁴⁺ salt/Y zeolite samples were carried out at 250 °C, 350 °C and 400 °C, respectively, in O₂ gas phase at a flow rate of 200 ml/min, and then followed by reduction in a 5% H₂/95% N₂ gas mixture at 360 °C or 400 °C to produce a fine distribution of Pd particles on zeolite.

Figure 1 illustrates the Pd migration pathways through zeolite structure under calcinations and reduction temperature control. The ion exchange process is to locate [Pd(NH₃)₄]⁴⁺ ligand on zeolite exterior surface and in supercages. The ammonia ligands start to decompose at 250 °C [13] and produce [Pd(NH₃)₂]²⁺ (see formula 1 in Fig. 1). The Pd²⁺ ions can migrate through zeolite channels and then further extend to zeolite small cages, i.e. sodalite cages and hexagonal prisms at calcinations temperature of 300 °C (see formula 2 in Fig. 1). Continuous increase of calcinations and reduction temperatures up to 350 °C and 360 °C respectively might force Pd atoms to migrate back into zeolite sodalite cages or zeolite supercages (see formula 3 in Fig. 1). Pd atoms can move back from zeolite small cages to supercages and zeolite external surface at calcinations and reduction temperature of 400 °C (see formula 3 in Fig. 1).

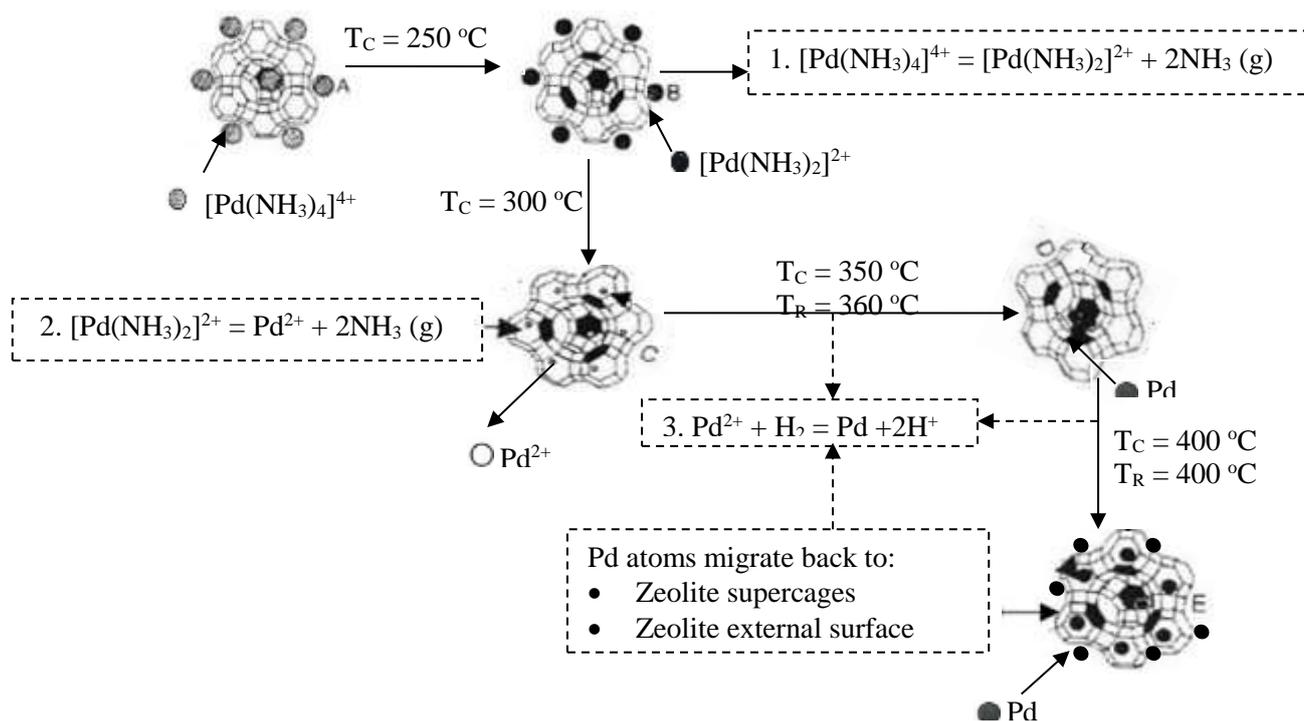
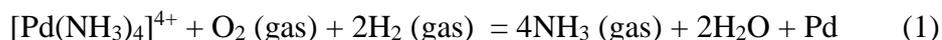


Figure 1. A schematic diagram depicts Pd migration controlled by calcinations (denoted as ‘C’) and reduction (denoted as ‘R’) process.

The following equation (1) prescribes an overall chemical reaction of Pd ligand by calcinations and reduction process:



2.3 Electrochemical cell and cyclic voltammetry (CV)

A typical 9 cm² electrode sheet was made of 1.5 wt% Pd loading on zeolite electrocatalyst mixed with XC-72R carbon powder (untreated) mixture and 15 wt% Nafion[®] solution (from Aldrich) as a binding agent. A solidification process was then taken place by hot pressing the resultant paste on carbon paper (E-TEK TGHP-90). A 2.5 cm² diameter disc electrode was trimmed for a standard CV measurement that was performed using a glass-jacket

electrochemical cell, consisting of a zeolite supported Pd working electrode, Hg/Hg₂SO₄ Mercury/Mercuries Sulphate (MMS) reference electrode and a Pd gauze counter electrode. The electrolyte was 2.5 mol dm⁻³ sulfuric acid (H₂SO₄) solution and the CV was measured at a scan rate of 1 mV s⁻¹ in a potential region of -0.675 V to 0.4 V for 1.5 wt% Pd loading on zeolite samples calcined at 250 °C and reduced at 360 °C (denoted as 15Pdanc25r36), calcined at 350 °C and reduced at 360 °C (denoted as 15Pdanc35r36) and calcined and reduced at 400 °C (denoted as 15Pdancr40), respectively, in which no electrolyte decomposition was detected.

2.4 The Ex-situ Extended X-ray Adsorption Fine Structure (EXAFS) Measurement

Pd/Y zeolite electrocatalysts were characterized by the ex-situ extended X-ray adsorption fine structure (EXAFS) measurement technique using a Synchrotron Radiation Source (SRS) at STFC Daresbury National Laboratory [7]. The wiggler beam line was operated at 2 GeV and 100 mA. Those high order harmonics that might affect the amplitude of EXAFS were removed using a double-crystal Si220 monochromator. The EXAFS data was then collected at the Pd K edge using a pellet placed in air. The pellet was made by 1.5 wt% Pd/Y zeolite catalyst powder mixed with boron nitride to give a Pd loading of 0.73 mg cm⁻². A Pd foil was used as a reference sample for EXAFS data collection. The resultant EXAFS data was later analyzed using in-house software EXCURV 98, where the inter-atomic distance, atom number and the type of backscattering neighbors were determined by a method proposed by Abruna [24].

3. Results and Discussion

3.1 The ex-situ extended X-ray adsorption fine structure (EXAFS) measurement

Pd particle size and distribution for 1.5 wt% Pd loading on zeolite were investigated using the ex-situ EXAFS technique. Figure 2, figure 3 and figure 4 display the EXAFS Chi data and

their Fourier transform fitting results using the ex-situ EXAFS measurement. The EXAFS data were collected at the Pd K edge in air. The fitted data are presented in tables 1, 2 and 3. The Chi data for catalyst samples 15Pdanc25r36 and 15Pdancr40 were valid over the range of wave vector from 2 \AA^{-1} to 18 \AA^{-1} , and for catalyst sample 15Pdanc35r36, Chi data were up to 17 \AA^{-1} . A data scattering was observed in Chi spectrum at a slightly high noise level after 14 \AA^{-1} . The EXAFS data was subsequently fitted up to 3 shells with respect to a goodness fitting value (R_{exafs}) of 40% representing an acceptable data fitting quality. The curve fitting (shown in dotted symbols) matches well against the raw data (shown in solid lines). The shells 1 & 2 Pd-Pd coordination numbers were predicted at 5.2 and 2.05 for catalyst 15Pdanc25r36 (table 1), 4.25 and 3.22 for catalyst 15Pdanc35r36 (table 2), and 5.77 and 2.22 for catalyst 15Pdancr40 (table 3), respectively. The Pd-Pd binding distance was determined at 2.74 \AA , indicating that Pd atom had metallic property in nature. A Pd-O shell was further added to improve data fitting quality for all three catalyst samples. The Pd-O distance was predicted at 2.01 \AA for catalyst samples 15Pdanc35r36 (table 2) and 15Pdancr40 (table 3) and 2.02 \AA for catalyst sample 15Pdanc25r36 (table 1), respectively. The presence of O neighboring atom indicates the oxidation of Pd to form Pd-O shell.

The Pd cluster size and geometry might be estimated using the first (or nearest) neighbor mean coordination number \overline{N}_1 of atoms in a cluster using method suggested by Benfield [25]. An icosahedra cluster is adopted to estimate the Pd cluster edge length using \overline{N}_1 , due to Pd metallic feature in nature. An average theoretical first (or nearest) neighbor mean coordination number and a total number of atoms in a cluster was estimated at 6.45 and 13, using equations (2) and (3) [25]. The Pd-Pd first shell coordination number estimated by EXAFS is slightly lower than those given by Benfield theory [25]. On average the Pd atoms are estimated at 11, 10 and 12, respectively based on EXAFS data analysis, which results in the following order of Pd particle size in radius: 15Pdancr40 at 1.0 nm > 15Pdanc25r36 at 0.9 nm > 15Pdanc35r36 at

0.8 nm. The EXAFS results imply that Pd particle size increases with the increase of calcinations and reduction temperatures. The larger particle size is usually correspondent to a low Pd active surface area based on per gram of metal. Thus, the expected electrochemical activity of charger transfer under the features of the hydride peak would be in an order of $15\text{Pdancr40} < 15\text{Pdanc25r36} < 15\text{Pdanc35r36}$, only if the calcinations and reduction temperature has solely effect on the particle size. However, this is not a type of trend being observed in CV measurement by determining charger transfer in hydride region. Therefore, the location of the metal particle on zeolite must also be taken into a consideration.

Unlike Pt, Pd is more capable to mobile from zeolite small cages to supercages and zeolite external surface under high calcinations and reduction temperature, due to its 'd' electron configuration [26] that can be altered with the change of Pd electronic and geometric structure [27]. The valence electrons on Pd are in the form of 4d orbital shape, whilst they are in the 5d orbital shape on Pt. Pd^{2+} has a greater charge density than Pt^{2+} . This will lead to the increase of the mobility of Pd. By controlling the calcinations and reduction temperature, Pd particle can be enriched in the zeolite supercages and zeolite external surface. Pd particles in these locations are more accessible to the protons in the electrolyte solution. In addition, Pd particles on the zeolite external surface might have a better electronic contact with the carbon support [7].

$$\overline{N_1}(i \cos) = 6 \left[(m-1)(2m^2 - 25m + 12) \right] / \left[(2m-1)(5m^2 - 5m + 3) \right] \quad (2)$$

$$N_{total} = (1/3)(2m-1)(5m^2 - 5m + 3) \quad (3)$$

Table 1. The ex-situ EXAFS data fitting results for data collected at the Pd K edge at ambient room temperature for catalyst sample 15Pdanc25r36

15Pdanc25r36	N	$2\sigma^2/\text{\AA}^2$	R/ \AA
Shell 1 O	0.95 ± 0.09	0.0049 ± 0.0012	2.02 ± 0.006
Shell 2 Pd	5.20 ± 0.62	0.0141 ± 0.0003	2.74 ± 0.003
Shell 3 Pd	2.05 ± 0.54	0.0125 ± 0.0032	3.41 ± 0.017

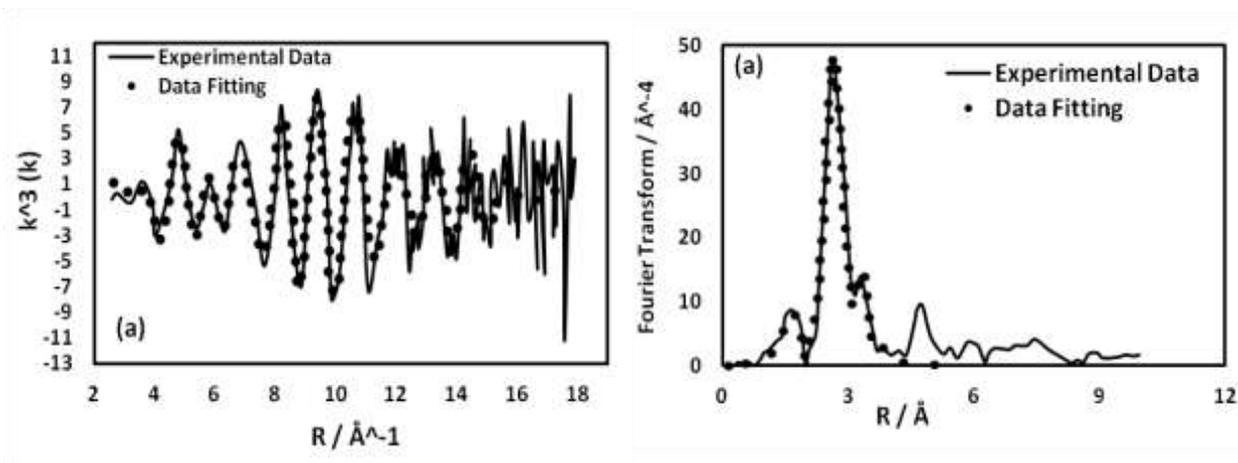


Figure 2. The ex-situ EXAFS Chi data and Fourier transforms and their fitting results for catalyst sample 15Pdanc25r36, phase corrected. The data were collected at ambient room temperature in air. The solid lines represent the experimental data, and the dotted symbols represent the fitted data.

Table 2. The ex-situ EXAFS data fitting results for data collected at the Pd K edge at ambient room temperature for catalyst sample 15Pdanc35r36

15Pdanc35r36	N	$2\sigma^2/\text{\AA}^2$	R/ \AA
Shell 1 O	1.45±0.25	0.0063±0.0023	2.02±0.006
Shell 2 Pd	4.25±0.28	0.0135±0.0006	2.74±0.003
Shell 3 Pd	3.22±0.57	0.0147±0.0016	3.42±0.010

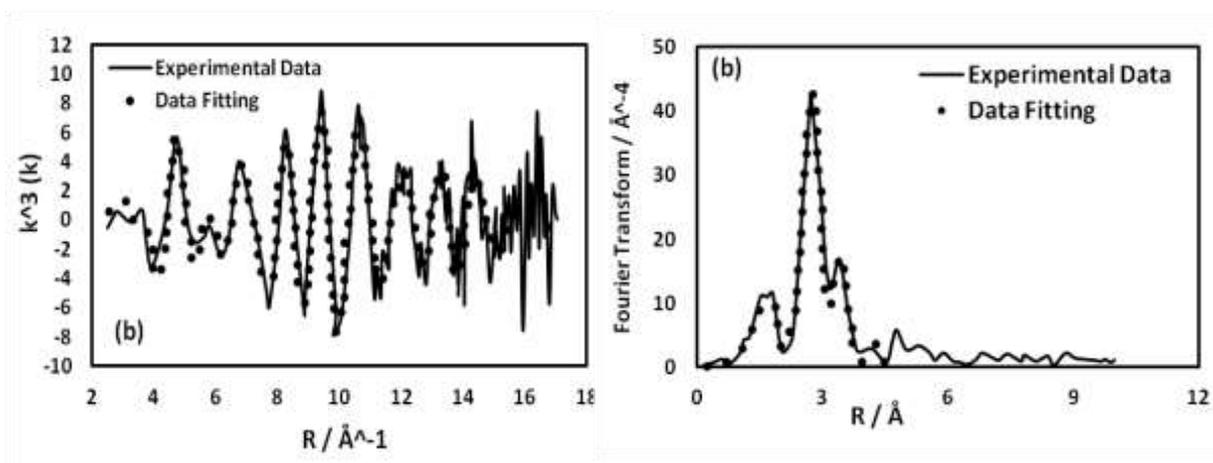


Figure 3. The ex-situ EXAFS Chi data and Fourier transforms and their fitting results for catalyst sample 15Pdanc35r36, phase corrected. The data were collected at ambient room temperature in air. The solid lines represent the experimental data, and the dotted symbols represent the fitted data.

Table 3. The ex-situ EXAFS data fitting results for data collected at the Pd K edge at ambient room temperature for catalyst sample 15Pdancr40

15Pdancr40	N	$2\sigma^2/\text{\AA}^2$	R/ \AA
Shell 1 O	1.10 ± 0.30	0.0072 ± 0.0038	2.01 ± 0.018
Shell 2 Pd	5.77 ± 0.33	0.0141 ± 0.0006	2.74 ± 0.005
Shell 3 Pd	2.22 ± 1.01	0.0200 ± 0.0053	3.39 ± 0.013

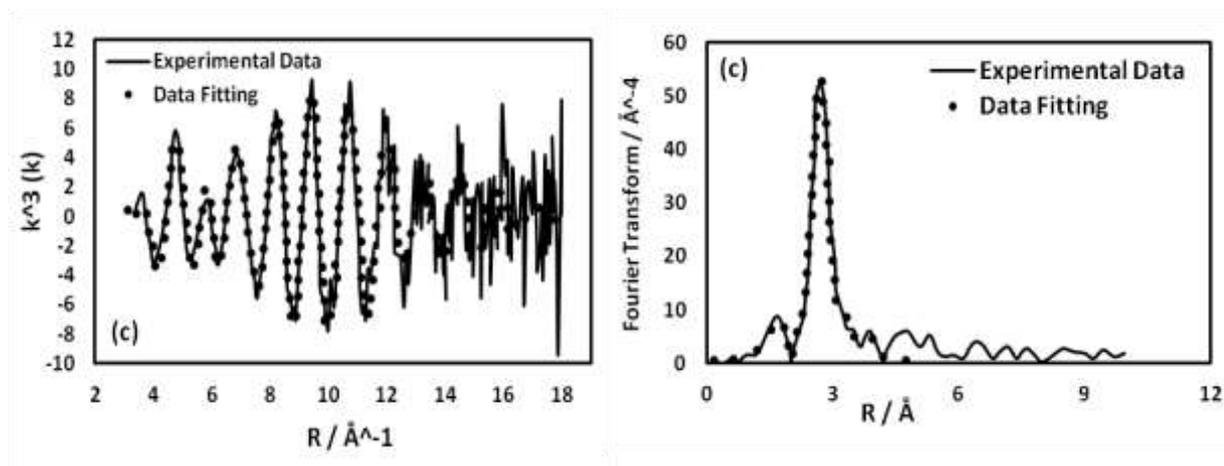


Figure 4. The ex-situ EXAFS Chi data and Fourier transforms and their fitting results for catalyst sample 15Pdancr40, phase corrected. The data were collected at ambient room temperature in air. The solid lines represent the experimental data, and the dotted symbols represent the fitted data.

3.2 Cyclic Voltammetry (CV) Measurement

The total accessible electrochemically active Pd surface area of electrode was measured in 2.5 mol dm^{-3} sulfuric acid (H_2SO_4) solution, corresponding to catalyst samples 15Pdanc25r36, 15Pdanc35r36 and 15Pdancr40. Figure 5 displays a diagram of CV measurement using 1.5 wt%

Pd on XC-72R Nafion[®] bound working electrode made by a commercial catalyst of 40 wt% Pd on XC-72R carbon mixed with extra XC-72R carbon powder (denoted as 15PdXC-72R) [28]. The two distinguished regions are identified and they are likely associated with the hydrogen adsorption, evaluation, hydrogen desorption and the re-oxidation peaks, respectively. This may be contributed by the morphology of Pd electrocatalyst [29]. A small peak 'A' is detected at -0.59 V, which is correspondent to hydrogen adsorption on Pd active site. The hydrogen evolution (i.e. the peak 'B') is observed at -0.689 V, and a double layer (DL) current density is measured between -0.18 V and -0.420 V in the cathodic sweep and from 0.18 V to -0.4 V in the anodic sweeps, respectively. The hydrogen re-oxidation represented by the peak 'C' is determined at -0.626 V on the anodic sweep and the second peak 'D' at -0.564 V is attributed to hydrogen desorption. The increase of current density from 0.226 V to 0.4 V is due to the formation of Pd oxide (PdO) on Pd surface. A peak 'E' predicted at 0.0656 V interprets the stripping of O formed on Pd surface.

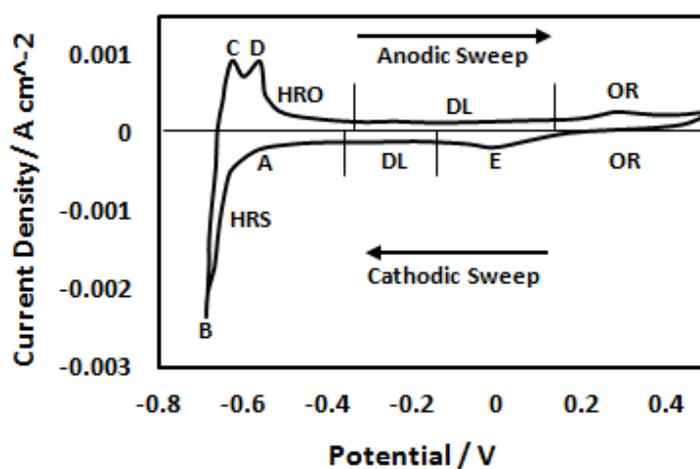


Figure 5. Steady State Cyclic Voltammogram (CV) measurement using 1.5 wt% Pd on XC-72R carbon powder Nafion[®] bound working electrode made by a commercial 40 wt% Pd on XC-72R carbon electrocatalyst mixed with extra carbon powder (denoted as 15PdXC-72R),

scanned at 1 mV s^{-1} in $2.5 \text{ mol dm}^{-3} \text{ H}_2\text{SO}_4$ at the potential region of -0.675 V to 0.4 V vs MMS reference electrode [28].

The electrochemical activity of Pd on zeolite was determined using a Nafion[®] bound Pd zeolite electrode fabricated by 1.5 wt% Pd loading on zeolite with extra XC-72R carbon power. Figure 6 has shown CV comparisons of electrochemical performance for 1.5 wt% Pd zeolite samples made by calcinations and reduction process at different temperatures. The strong current density change was detected using electrode made by catalyst sample 15Pdancr40 calcined and reduced at $400 \text{ }^\circ\text{C}$ in hydride region, where two distinguished peaks for weakly or strongly hydrogen desorption were captured at -0.607 V and -0.631 V in the anodic sweep. The peak tendency measured for catalyst sample 15Pancr40 is very identical to that measured by catalyst sample 15PdXC-72R. The current density of hydrogen desorption has seen a smaller increase for catalyst sample 15Pdanc35r36 than catalyst 15Pdanc25r36 in hydrogen adsorption and desorption regions, respectively, where a single hydrogen desorption peak is observed for both catalysts. There is no second hydrogen desorption peak captured in the hydride region, which implies these two hydrogen desorption peaks having a very similar energy level on Pd active sites. The current density increase in the hydride region follows an order of $15\text{Pdancr}40 > 15\text{Pdanc}35\text{r}36 > 15\text{Pdanc}25\text{r}36 > 15\text{PdXC-}72\text{R}$. A similar tendency was detected in the oxygen oxidation and reduction region for all the four samples. The CV measurement indicates that the calcinations and reduction temperature increasing up to $400 \text{ }^\circ\text{C}$ enables the improvement of Pd electrocatalytic activity in an electrochemical environment. The similar peak areas captured in the hydride region for catalyst samples 15Pdanc25r36 and 15Pdanc35r36 indicate that Pd particle has a very similar location in the structure of these two samples. The calcinations and reduction temperature has played a critical role to mobile the Pd atoms through the zeolite channel and cages. The Pd atoms can also migrate out from zeolite

hexagonal prisms to sodalite cages at the calcinations temperature of 350 °C. The mobility of Pd particles from zeolite small cages to the supercages and zeolite external surface are more pronounced by high calcinations and reduction temperature at 400 °C. This has resulted in the increase of the Pd active surface areas. Furthermore, the restoration of zeolite active sites (i.e. Brønsted acidic sites) under hydrogen reduction process might have enhanced the migration of Pd atoms between the zeolite channels and cages [16].

Previously, the in-situ EXAFS data analysis to determine Pd particle structure in the hydride region via electrode made by catalyst samples of 15Pdanc35r36 has depicted a Pd-Pd distance of 2.83 Å at -0.675 V and 2.80 Å at -0.65 V which has seen a significant increase whereas compared to a metallic Pd-Pd distance of 2.74 Å (see table 4) [28]. This increase corresponding to Pd distance might be due to hydrogen adsorbed on Pd surface or forming Pd hydride phases such as β -PdH or α -PdH on Pd particle through ‘spillover’ of hydrogen atom on zeolite channel or surface conductance. The β -PdH associated with a direct hydrogen adsorption reaction between electrode surface and sub-surface might be decomposed to α -PdH, due to the migration of hydrogen atom from a nearer surface of Pd hydride phase to the center of Pd particle. The Pd-Pd atom binding distance was predicted at 2.74 Å at -0.5 V and 2.73 Å at 0.5 V which implies the Pd particle presenting more metallic features in nature. The Pd particle size tends to be large at more negative potential in the hydride region as indicated by the Pd first shell coordination number (see table 4), and this might be directly associated with the formation of β -PdH and α -PdH hydride phase, resulting in a long Pd-Pd distance in comparison to the Pd-Pd distance predicted for same catalyst samples in atmosphere condition.

Table 4. The in-situ EXAFS data fitted at the Pd K edge in hydride region at the potential of -0.675 V, -0.65 V, -0.5 V and + 0.5 V via MMS reference electrode [28].

Shell 1_Pd / Potential (V)	N Pd coordination number	R/ Å Pd-Pd distance
-0.675 V	7.86±0.46	2.83±0.005
-0.650 V	7.09±0.31	2.80±0.004
-0.500 V	5.65±0.61	2.74±0.005
+0.500 V	6.18±0.52	2.73±0.008

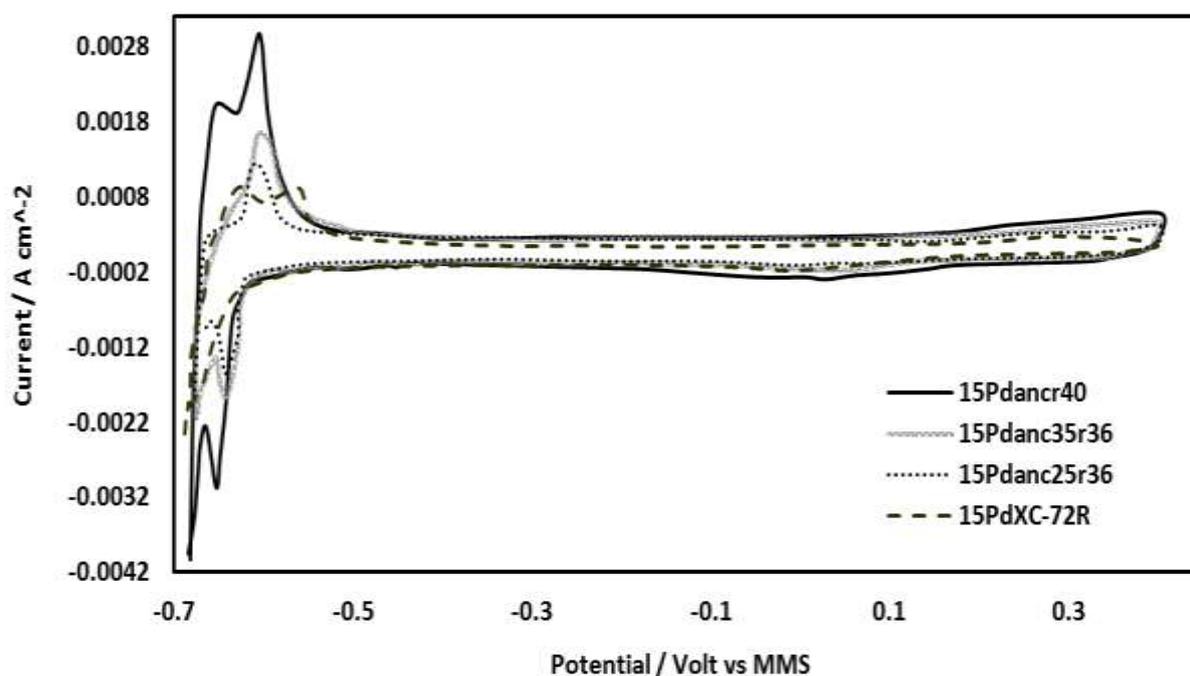


Figure 6. Comparison of CV measurements to determine the catalyst performance of 15Pdancr40 – calcined and reduced at 400 °C; 15Pdanc35r36 – calcined at 350 °C and reduced at 360 °C; 15Pdanc25r36 – calcined at 250 °C and reduced at 360 °C, respectively. The CV was measured at scan rate of 1 mV s⁻¹ in 2.5 mol dm⁻³ H₂SO₄ at the potential region of -0.675 V to 0.4 V vs MMS reference electrode.

4. Conclusion

The influence of calcinations and reduction temperature on Pd particle size has been investigated by the ex-situ extended X-ray adsorption fine structure (EXAFS) and Cycle Voltammetry (CV) techniques. The enhancement of electrochemical activity associated with the charger transfer on Pd active surface in the hydride region increases with calcinations and reduction temperature up to 400 °C. No major discrepancy of current change with respect to hydrogen adsorption and desorption peaks was determined for catalyst samples calcined at 250 °C and reduced at 360 °C (i.e. 15Pdanc25r36) or calcined at 350 °C and reduced at 360 °C (i.e. 15Pdanc35r36), respectively. The Pd particle size follows an order of 15Pdancr40 (1.0 nm) > 15Pdanc25r36 (0.9 nm) > 15Pdanc35r36 (0.8 nm). A high dispersion of small metal particle size is usually correspondent to a high catalyst activity. However, this tendency is not reflected by CV measurements observed for electrochemical activity of charger transfer in the hydride region, in which the electrocatalytic activity of catalysts is arranged in an order of 15Pdancr40 > 15Pdanc35r36 > 15Pdanc25r36. The Pd location might have played an important role to increase the catalytic activity of Pd particle. For catalyst 15Pdancr40, the majority of Pd atoms are resident at zeolite supercages and/or on the zeolite external surfaces. For those samples made at the calcinations and reduction temperature below 360 °C, the Pd atoms are most likely to remain at zeolite small cages such as sodalite cages or hexagonal prisms. The mobility of Pd atoms can be facilitated by atomic hydrogen disassociated by H₂ molecule or the restoration of zeolite active sites (i.e. Brønsted acidic sites) during a chemically H₂ reduction process.

5. References

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